Hoffmann- Bromamide Rearrangement

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Introduction:

The reaction involves the conversion of an amide into a primary amine with one carbon less, by the action of alkaline hypohalite (NaOH solution + Br_2 or Cl_2). The overall reaction is:

$$RCONH_2 + Br_2 + 4NaOH \longrightarrow RNH_2 + 2NaBr + Na_2CO_3 + 2H_2O$$

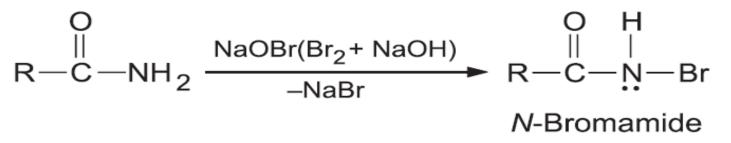
where R may be aliphatic, aromatic or heterocyclic.

Mechanism:

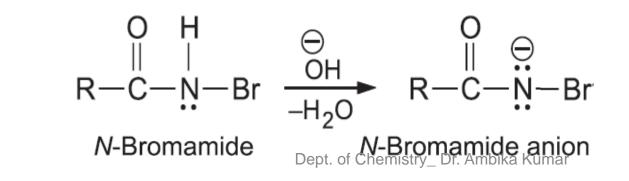
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The mechanism has been suggested on the basis of the intermediates isolated during the course of reaction.

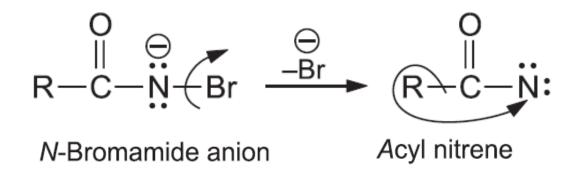
In the first step N-bromamide is formed by the usual reaction of hypobromite on primary amines:



This follows the removal of the acidic hydrogen from the N-atom by the base to form N-bromamide anion in the second step:



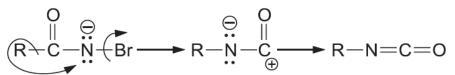
In the third step (rate determining step) loss of bromine results in the formation of the highly reactive acyl nitrene.



Note: The possibility of the formation of acylnitrene, however, has been rejected, for if it is formed it would react with water to give hydroxamic acid, which has not been detected.

$$\begin{array}{c} O \\ \parallel \\ R - C - N \\ \vdots \\ + \\ H_2 O \longrightarrow R - C - NHOH (Hydroxamic acid) \end{array}$$

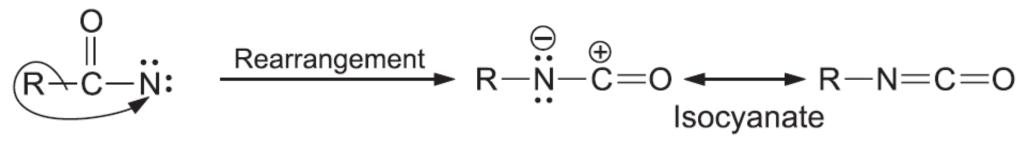
Hence, the loss of Br and the migration of R in *N*-bromamide anion is suggested to be concerted (since no cross over products are formed and migrating R retains its chirality – if molecule is asymmetric).



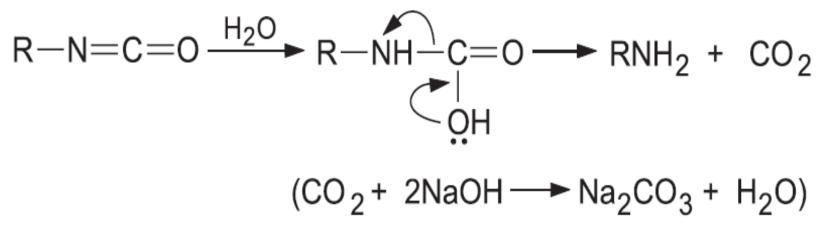
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Thus in the fourth step, migration of R to the electron-deficient nitrogen gives the isocyanate.



Formed isocyanate reacts with water to give carbamic acid which spontaneously eliminates carbon dioxide to yield the amine.



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Applications

Formation of amines It is a very useful method for the preparation of amines.
 (a) Acids and amides to amines.

(i) RCOOH
$$\frac{1. \text{ NH}_4\text{OH}}{2.\Delta}$$
 RCONH₂ $\xrightarrow{\text{Br}_2 + \text{ KOH}}$ RNH₂
(ii) Me₃CCH₂CONH₂ $\xrightarrow{\text{NaOBr}}$ Me₃CCH₂NH₂
 β, β -Dimethyl butyramide Neopentyl amine

(b) Higher amides (with more than eight carbons) give nitriles which can be reduced to amines.

$$\mathsf{RCH}_2\mathsf{CONH}_2 \xrightarrow{\mathsf{Br}_2 + \mathsf{KOH}} \mathsf{RCH}_2\mathsf{CN} \xrightarrow{\mathsf{Na} + \mathsf{Alcohol}} \mathsf{RCH}_2\mathsf{CH}_2\mathsf{NH}_2$$

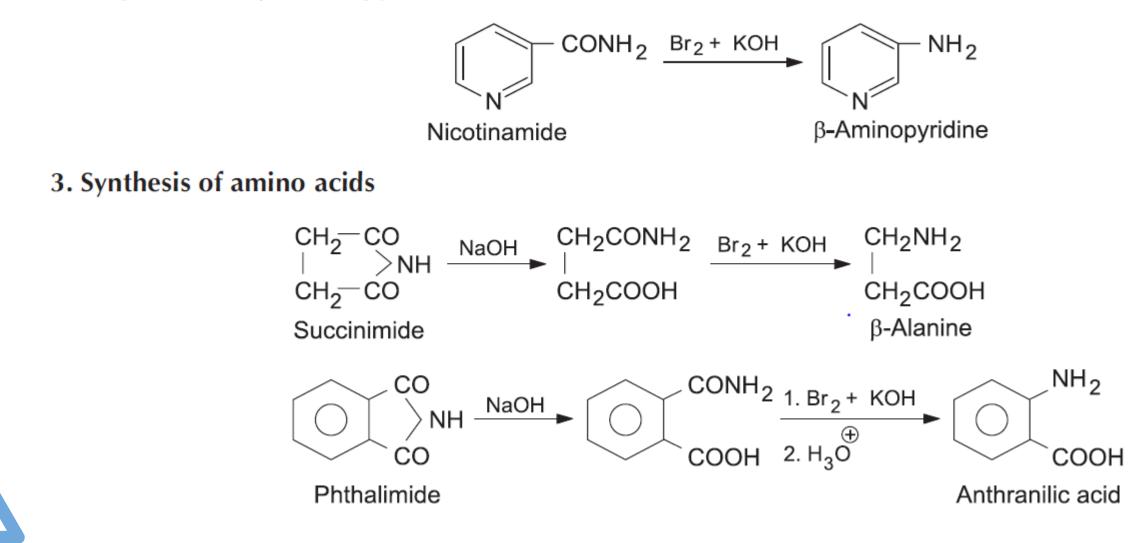
(c) Urea gives the valuable reagent hydrazine.

H₂NCONH₂
$$\xrightarrow{\text{Br}_2 + \text{KOH}}$$
 NH₂NH₂
Urea

The yield of amines is poor with β , γ -unsaturated acids.

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2. Preparation of β-amino pyridine from nicotinamide



4. Preparation of aldehydes from hydroxyacid amides

(i)
$$C_6H_5$$
-CH(OH)-CONH₂ $\xrightarrow{Br_2 + KOH} \begin{bmatrix} C_6H_5 - CH + NH_2 \end{bmatrix} \xrightarrow{O} C_6H_5 - C - H + NH_3$
Mandelic acid amide

(ii) Methanolic α, β-unsaturated acid amides on treatment with NaOCI give urethane which on hydrolysis with HCI gives a good yield of aldehyde.

$$R-CH=CH-CO-NH_{2} \xrightarrow{NaOCl} R-CH=CH-NHCOOCH_{3} \xrightarrow{HCI} R-CHO$$