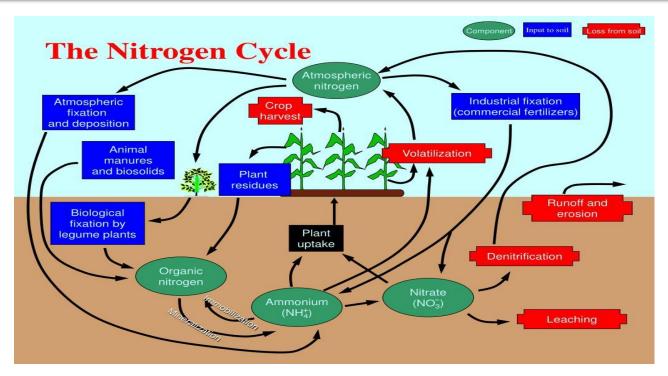


Bhagalpur National College, Bhagalpur

(A Constituent unit of Tilka Manjhi Bhagalpur University, Bhagalpur)

PPT Presentation for B.Sc. III- Environmental Pollution



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The **nitrogen cycle** is the process by which **nitrogen** is converted between its various chemical forms. This transformation can be carried out through both biological and physical processes.

Forms of Nitrogen :

a) organic nitrogen as-

- ammonium (NH₄⁺),
- nitrite (NO₂⁻),
- nitrate (NO₃⁻),
- nitrous oxide (N₂O),
- nitric oxide (NO) or

b) inorganic nitrogen as $nitrogen gas(N_2)$.

Nitrogen cycle consists of the following steps-

- 1. Nitrogen Fixation
- 2. Nitrogen assimilation
- 3. Ammonification
- 4. Nitrification and
- 5. Denitrification
- 6. Sedimentation

1. Nitrogen fixation :

The conversion of free nitrogen of atmosphere into the biologically acceptable form or nitrogenous compounds.

There are following ways to convert N₂ into more chemically reactive forms:

- a) Biological Nitrogen fixation
- b) Physiocochemical nitrogen fixation
- c) Industrial nitrogen fixation

a) **Biological Nitrogen fixation :**

some symbiotic bacteria , blue-green algae and some free-living bacteria are able to fix nitrogen as organic nitrogen.

e.g-

symbiotic bacteria: Rhizobiumsymbiotic blue-green algae : species of Nostoc, Anabaena , etcfree-living bacteria: Azotobacter, Clostridium, Derxia,

Rhodospirillium, etc

b) Physiocochemical or Non-biological nitrogen fixation :

In this process, atmospheric nitrogen combines with oxygen (as ozone) during lightning or electrical discharges in the clouds and produces different nitrogen oxides :

$$N_{2} + 2(O) \xrightarrow{Electric} 2NO$$
$$2NO + 2(O) \xrightarrow{Discharge} 2NO_{2}$$
$$2NO_{2} + (O) \rightarrow N_{2}O_{5}$$

The nitrogen oxides get dissolved in rain water and on reaching earth surface they react with mineral compounds to form nitrates and other nitrogenous compounds :

> $N_2O_5 + H_2O \rightarrow 2HNO_3$ $2HNO_3 + CaCO_3 \rightarrow Ca(NO_3)_2 + CO_2 + H_2O$

c) Industrial nitrogen fixation :

Under great pressure, at a temperature of $600^{\circ}C$ and with the use of an iron catalyst, hydrogen and atmospheric nitrogen can be combined to form ammonia (NH₃) in the Haber-Bosch process.

2. Nitrogen assimilation :

In this process, Inorganic nitrogen in the form of nitrates, nitrites, and ammonia is absorbed by the green plants via their roots and then it is converted into nitrogenous organic compounds.

Nitrates are first converted into ammonia which combines with organic acids to form aminoacids . Aminoacids are used in the systhesis of proteins, enzymes, chlorophylls, nucleic acids, etc.

3. Ammonification :

It is the process of releasing ammonia by certain microorganisms utilizing organic compounds derived from the dead organic remains of plants and animals and excreata of animals .

The microorganisms especially involved are-

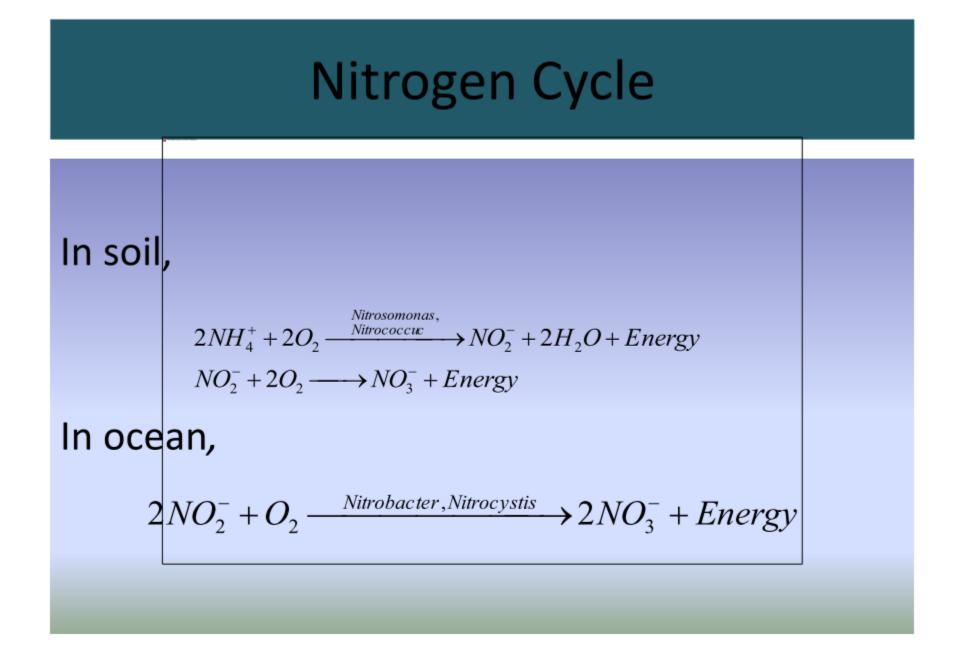
actinomycetes, and bacilli (Bacillus ramosus, B. vulgaris, B. mesenterilus)

4. Nitrification :

Nitrification is a process of enzymatic oxidation of ammonia to nitrate by certain microorganisms in soil and ocean.

Nitrosomonas ammonia to nitrites (NO2-

Nitrobacter oxidation of the nitrites into nitrates (NO₃⁻).



5. Denitrification :

Denitrification is the reduction of nitrates back into the largely inert nitrogen gas (N_2) .

Some denitrifying bacteria are-

Thiobacillus denitrificants Micrococcus denitrificants Pseudomonas aeruginosa

$2NO_3^- \rightarrow 2NO_2^- \rightarrow 2NO \rightarrow N_2O \rightarrow N_2$

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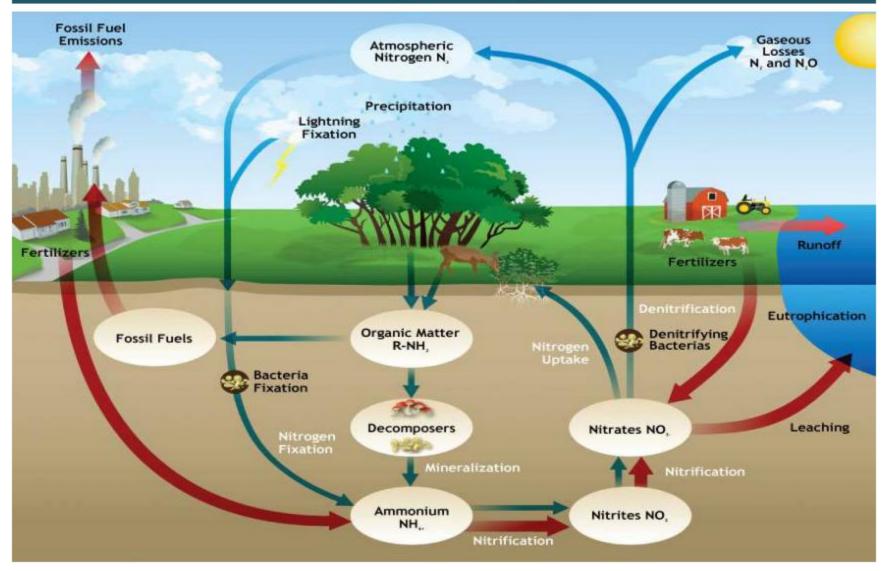
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6. Sedimentation :

Sometimes, nitrates of soil are locked up in the rocks while they are washed down to the sea or leached deeply into the earth along with percolating water. This phenomena is known as **sedimentation**.



Thank You All



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